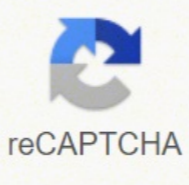




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Underlined words and phrases are to be filled in by students on the Note-taking Worksheet.

Section 1 Stability in Bonding

- A. Some elements combine chemically and no longer have the same properties they did before forming a compound.
- B. A(n) chemical formula is composed of symbols and subscripts indicating the number of atoms of an element in a compound.
- C. Atoms form compounds when the compound is more stable than the separate atoms.
 1. Noble gases are more chemically stable than other elements because they have a complete outer energy level.
 2. Elements that do not have full outer energy levels are more stable in compounds.
 3. Atoms can lose, gain, or share electrons to get a stable outer energy level.
 4. A(n) chemical bond is the force that holds atoms together in a compound.

DISCUSSION QUESTION:
Why do elements form compounds? To become more chemically stable by getting a complete outer energy level

Section 2 Types of Bonds

- A. A(n) ion is a charged particle because it has more or fewer electrons than protons.
 1. When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge.
 2. When an atom gains an electron, it becomes a negatively charged ion.
- B. An ionic bond is held together by the ionic bond—the force of attraction between opposite charges of the ions.
 1. The result of this bond is a(n) neutral compound.
 2. The sum of the charges on the ions is zero.
- C. Molecules are neutral particles formed as a result of sharing electrons.
 1. A covalent bond is the force of attraction between atoms sharing electrons.
 2. Atoms can form double or triple bonds depending on whether they share two or three pairs of electrons.

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T2 Chemical Bonds

Chemistry: Chapter 1 & 2 Review

Name: _____

1. A mixture (is not) a chemical combining of substances.
2. In a compound the (atoms) molecules (are) physically (combined) so that the elements that make up the compound (retain) their identities and (do) (do not) take on new set of properties.
3. The smallest identifiable unit of a compound is (an) MOLECULE, which is made up of two or more atoms, which are chemically bonded.
4. True or False? A mixture is always made up of a combination of elements. COULD BE A COMBINATION OF COMPOUNDS
5. In a mixture, the substances (can) (retain) their identities.
6. In a mixture, the substances involved (can) (not) be separated by a simple physical process.
7. In a compound, the elements involved (can) (not) be separated by a simple physical process because the elements are (physically combined) (chemically bonded).
8. The smallest identifiable unit of an element is a(n) ATOM.
9. Explain how to separate the sugar and water in a solution of sugar and water. BOIL (OR EVAPORATE) THE WATER OFF, LEAVING THE SUGAR BEHIND.
10. How would you separate a mixture of alcohol and water? DISTILLATION - HEAT MIXTURE TO THE 1° BOILING POINT (ALCOHOL) Collect Vapors --> Condensation --> Liquid
11. How would you separate sand and water? FILTER - WATER WOULD PASS THRU LEAVING THE SAND ON THE FILTER
12. Classify the following as pure substances or as mixtures:

Gasoline	<u>1</u>	<u>M</u>	Sugar	<u>2</u>	<u>or</u>	<u>3</u>
Mercury	<u>3</u>		Tap water	<u>4</u>		Oxygen
Distilled Water	<u>5</u>		Blood	<u>6</u>		Nitrogen Dioxide

13. What is the difference between Extensive and Intensive Properties?

14. Why are "intensive properties" so important?

Small text at the bottom left corner, likely a page number or reference code.

WRITING FORMULAS FROM NAMES

Name Key

Write the formulas of the following compounds.

- ammonium phosphate $(\text{NH}_4)_3\text{PO}_4$
- iron (II) oxide FeO
- iron (III) oxide Fe_2O_3
- carbon monoxide CO
- calcium chloride CaCl_2
- potassium nitrate KNO_3
- magnesium hydroxide $\text{Mg}(\text{OH})_2$
- aluminum sulfate $\text{Al}_2(\text{SO}_4)_3$
- copper (II) sulfate CuSO_4
- lead (IV) chromate $\text{Pb}_2(\text{CrO}_4)_4 \Rightarrow \text{Pb}_6(\text{CrO}_4)_2$
- diphosphorus pentoxide P_2O_5
- potassium permanganate KMnO_4
- sodium hydrogen carbonate NaHCO_3
- zinc nitrate $\text{Zn}(\text{NO}_3)_2$
- aluminum sulfite $\text{Al}_2(\text{SO}_3)_3$

Chapter 5: Ionic Compounds Test

Name: _____

Score: _____/40

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- In many compounds, atoms of main-group elements form ions so that the number of electrons in the outermost energy levels of each ion is
 - 2.
 - 6.
 - 8.
 - 10.
- An anion
 - is an ion with a negative charge.
 - attracts ions with negative charges.
 - results when an alkaline-earth metal loses one of its two outermost electrons.
 - has more protons than electrons.
- When the octet rule is satisfied, the outermost _____ are filled.
 - d* and *f* orbitals
 - s* and *p* orbitals
 - s* and *d* orbitals
 - d* and *p* orbitals
- The elements of the _____ group satisfy the octet rule without forming compounds.
 - halide
 - noble gas
 - alkali metal
 - alkaline-earth metal
- Once an atom has full *s* and *p* orbitals in its outermost energy level,
 - it is highly reactive only with alkali metals.
 - it is highly reactive only with halogens.
 - it can be combined with most elements.
 - it has a stable octet and is unreactive.
- In a crystal of an ionic compound, each cation is surrounded by
 - molecules.
 - positive ions.
 - dipoles.
 - anions.
- The indium(II) ion and indium(III) ion
 - have the same charge.
 - are polyatomic ions.

Crispix has some of the properties of a compound. The individual components of a mixture can be physically separated from each other. Cereal does not have a constant composition; The ratio between raisins to bran flakes changes from the sample to the sample. By convention, no subscript is written when a molecule contains only one atom of an element. 1. These compounds are composed of positive and negative ions formed by add or subtracting electrons from neutral atoms and molecules. 4. The elements of atoms consist of atoms, the smaller particle that has one of the properties of element, John Dalton, in 1803, proposed a modern atom theory based on the following hypotheses. All the atoms of an element are identical. 5. 2. It can contain the minimum of 10% or 45%, zinc. Water is always 88.8% or 11.2% H by weight regardless of its source. Some of the differences between chemical compounds and mixtures of elements are illustrated by the following example using the raisin bran and "crispix". The notation of the shorthand for a compound describes the number of atoms of each element, which is indicated by an index written after the symbol for the element. When a mixture decomposes, the atoms are retrieved unchanged. It is easy to physically separate the two "elements" to choose the raisins, for example, and eat them separately. Compound fusion point (OC) Boiling point (OC) CR2O3 2266 4000 SRF2 1470 2489 CCL4 -22.9 76.6 CHC4 -22.9 76.6 CHC4 -22.9 76.6 CH3OH -97.8 64.7 Click here to check the response to the practice problem 2 formulas a molecule is the smallest particle Which has any of the compound properties. The semimetals can be found along the dividing line between metals and metals. 3. Characteristics of ionic and covalent compounds Covalent compounds contain positive and negative ions (NA + CL-) exist as sag sag o iduqil, idilos)s[lcan' alovat ad elas emoc heat SSILLOS 12O21H6G(ertuen Points H2O (L), CO2 (G)) Melting and hot co2 points (g)) Melting and hot points (ie often exist as liquid or gas at room temperature) strong force of attraction between relatively weak particles Attracting between the separate molecules in loaded the particles in water to give a solution that leads the electricity remain as the same molecule in the water and will not lead the electricity determining if a mixture is ionic or covalent calculates the difference between the electronegities of two Elements in a mixture and average of their electronegities and finds the intersection of these values on the figure shown below to determine if the mixture is ion or covalent or metallic. A handful of non-metal is grouped into the upper right corner of the periodic table. Practice problem 1: For each of the following compounds, predict whether you would expect that it is ionic or covalent. Atoms cannot be created or destroyed. The matter is composed of indivisible and indestructible atoms. Atoms of different elements have different weights and different chemical properties. A. (a) chromium (iii) oxide, cr2o3 (b) tetrachloride of carbon, ccl4 (c) methanol, ch3oh (d) strontium fluoride, srf2 click here to verify the answer to the practice problem 1 practice problem 2: use i Following data on proposing a way to distinguish between ionic and covalent compounds. The chemical compounds are very different from mixtures: the elements in a chemical compound can only be separated by destroying the mixture. The mixtures, such as the atmosphere, contain two or more relatively easy substances to separate. The formula for a molecule must be neutral. The raisins Bran has the following properties of a mixture. Therefore, water is H2O and carbon dioxide is CO2. The law of constant composition can be used to distinguish between compounds and mixtures of The compounds have a constant composition: Mixtures don't. do. It is an example of a mixture of two elements: copper and zinc. There is no way to separate the "elements" without breaking the bonds that hold them together. Hydrogen and oxygen, on the other hand, cannot be broken down into simpler substances. Most of the elements are metals, which are located to the left and bottom of the periodic table. Elements can be divided into three categories that have characteristic properties: metals, non-metals and semi-metals. Compounds Elements combine to form chemical compounds which are often divided into two categories. So it is the elementary chemicals, or simpler, the elements. Metals often react with non-metals to form ionic compounds. The notation for each element can be found on the periodic table of elements. Each element is represented by a unique symbol. Elements, compounds and mixtures Elements Any substance that contains only one type of atom is known as an element. The ratio between rice flakes and corn flakes is constant; It's 1:1 in every sample. Non-metals combine to form covalent compounds, which exist as neutral molecules. When writing the formula for an ionic compound, the charges on the ions must be balanced, the number of positive charges must be equal to the number of negative charges. The atoms of different elements combine into simple whole numbers to form compounds. Another difference between compounds and mixtures of elements is the ease with which the elements can be separated. Examples: CaCl2 balanced formula has 2 positive charges (1 calcium ion with charge +2) and 2 negative charges (2 chloride ions with charge -1) Al2 (SO4) 3 Balanced formula has 6 positive charges (2 aluminum ions with charge +3) and 6 negative charges (3 sulfate ions with charge -2) Mixtures Vs. Compounds The law of constant composition states that the mass ratio of the elements of a compound .otsopmoc .otsopmoc led etnof allad etnemtednepidni ,ossets ol erpmes "A Elements such as phosphorus (P4) or sulfur (S8) cannot be disrupted in simpler substances. Example: The water decomposes in a mixture of hydrogen and oxygen when an electric current passes through the liquid. liquid.

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